SCIENCE 9
HYDROLYSIS LAB

INTRODUCTION: When certain salts are dissolved in water, the resulting solutions are slightly acid. Others form solutions that show slightly basic properties. This is known as **Hydrolysis of Salts.**

In this lab you will test a number of salts to determine which type of salt it is and you will also determine which acid and base were used to make the salt.

DATA TABLE

SALT USED	LITMUS CHANGE RED	LITMUS CHANGE BLUE	TYPE OF SALT	ACID AND BASE USED
Sodium borate				
Copper (II) sulfate				
Sodium carbonate				
Sodium chloride				
Potassium nitrate				
Iron (III) chloride				
Lead ~~acetate~~ nitrate				
Aluminum sulfate				
Ammonia chlorate				
Sodium phosphate				

ACIDS

sulfuric H_2SO_4

hydrochloric HCl

nitric HNO_3

carbonic H_2CO_3

boric H_3BO_3

phosphoric H_3PO_4

acetic $HC_2H_3O_2$

tartaric $H_2C_4H_4O_6$

hydrofluoric HF

sulfurous H_2SO_3

BASES

sodium hydroxide $NaOH$

calcium hydroxide $Ca(OH)_2$

potassium hydroxide KOH

magnesium hydroxide $Mg(OH)_2$

aluminum hydroxide $Al(OH)_3$

ammonia-water NH_4OH

iron(III) hydroxide $Fe(OH)_3$

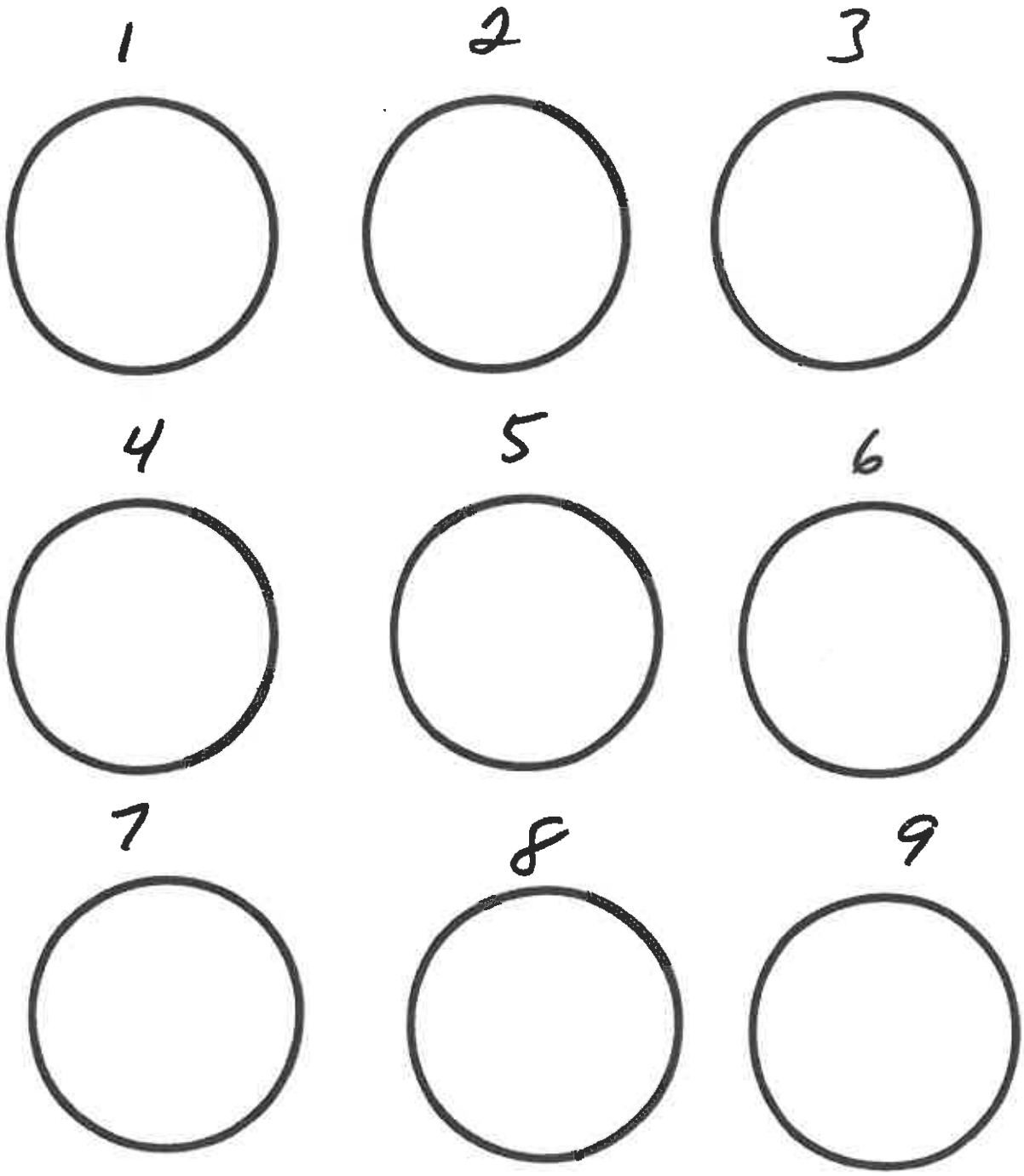
copper (II) hydroxide $Cu(OH)_2$

lithium hydroxide $LiOH$

lead hydroxide $Pb(OH)_2$

barium hydroxide $Ba(OH)_2$

zinc hydroxide $Zn(OH)_2$


1
2
3
4
5
6
7
8
9